

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 16

## STUDY GUIDE FOR CONTENT MASTERY

## Energy and Chemical Change

## Section 16.1 Energy

In your notebook, read about the nature of energy.

In the space at the left, write true if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- True* \_\_\_\_\_ 1. Energy is the ability to do work or produce heat.
- cannot be* \_\_\_\_\_ 2. The law of conservation of energy states that energy can be created and destroyed.
- True* \_\_\_\_\_ 3. Chemical potential energy is energy stored in a substance because of its composition.
- True* \_\_\_\_\_ 4. Heat is a form of energy that flows from a warmer object to a colder object.
- True* \_\_\_\_\_ 5. A calorie is the amount of energy required to raise the temperature of one gram of pure water by one degree Celsius.
- False* \_\_\_\_\_ 6. A calorie is the SI unit of heat and energy.
- True* \_\_\_\_\_ 7. The specific heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.

*True* \_\_\_\_\_ 8. Kinetic energy is energy of motion.

*both potential and kinetic energy* \_\_\_\_\_ 9. Chemical potential energy is a chemical reaction *contains only* potential energy.

*1000 calories* \_\_\_\_\_ 10. One nutritional Calorie is equal to 100 calories.

*True* \_\_\_\_\_ 11. One calorie equals 4.184 joules.

*True* \_\_\_\_\_ 12. When a fuel is burned, some of its chemical potential energy is lost as heat.

*multiply* \_\_\_\_\_ 13. To convert kilojoules to joules, divide the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following questions. Show all your work.

14. If the temperature of a 500-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/g·°C.

$$q = c \times m \times \Delta T$$

$$q = 4.184 \text{ J/g}\cdot^{\circ}\text{C} \times 05 \times 10^3 \text{ g} \times 2.00^{\circ}\text{C}$$

$$q = 4.184 \times 10^4 \text{ J} = 4.18 \text{ kJ}$$

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Chemistry: Matter and Change • Chapter 16

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## CHAPTER 16

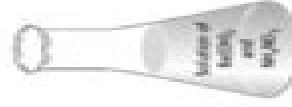
## STUDY GUIDE FOR CONTENT MASTERY

## Section 16.2 Heat in Chemical Reactions and Processes

In your notebook, read about measuring heat and about chemical energy and the universe.

For each item in Column A, write the letter of the matching item in Column B.

- | Column A   | Column B                       |
|--|--------------------------------|
| b. An insulated device used to measure the amount of heat absorbed or released during a chemical or physical process | a. system                      |
| c. The study of heat changes that accompany chemical reactions and phase changes                                     | b. calorimeter                 |
| d. The specific part of the universe that contains the reaction or process you wish to study                         | c. thermodynamics              |
| e. The change in enthalpy in a chemical reaction   | d. universe                    |
| f. A system plus its surroundings  | e. enthalpy                    |
| g. The heat content of a system at constant pressure   | f. enthalpy change of reaction |
| h. Everything in the universe except the system being studied  | g. surroundings                |



Use the illustration to answer the following questions.

9. A solution is studying the solution in the flask. What is the system?  
The system is the solution of NaOH and HNO<sub>3</sub>.
10. What are the surroundings?  
The surroundings include everything except the solution.
11. What is the universe?  
The universe is the solution plus the surroundings.

Study Guide for Content Mastery

Chemistry: Matter and Change • Chapter 16

# Energy And Chemical Change Study Guide Answer

**KJ Lindholm-Leary**



## **Energy And Chemical Change Study Guide Answer:**

*Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry* Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

**Energy: Forms & Changes Science Learning Guide** NewPath Learning, 2014-03-01 The Energy Forms Potential Energy Kinetic Energy Forms of Energy Energy Transformation Conservation of Energy Heat Sources of Energy Nonrenewable and Sources of Energy Renewable Aligned to Next Generation Science Standards NGSS and other state standards

**Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry** Katherine J. Denniston, Denniston Katherine, 2006

Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1996 This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker more realistic view of Botswana Sue Baker Publishing News

**Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition** John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter

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