Types of Chemical Reactions Lab

Chemical reactions can be categorized into five different types of reactions:

Synthesis Single Displacement Decomposition Double Displacement

Combustion

In this laboratory activity, you will experiment with various chemical reactions and classify them based on their reactants and products.

Materials

- Copper pellets (Cu)
- Magnesium strips Mg)
- Zinc metal (Zn)
- Silver nitrate (AgNO₃)

- Copper (II) sulfate (CuSO₄)
- 1M Hydrochloric acid (HCl)
- Hydrogen peroxide (H₂O₂)
- Potassium iodide (KI)

- Lead (II) nitrate (PbNO₃)
- Test tubes (3)
- Test tube holder
- · Bunsen burner

Procedure

Reaction A

- 1. Obtain a small number of copper (Cu) pellets. Place them into your test tube.
- 2. Carefully pour (or drip) silver nitrate (AgNO₃) into the test tube until it covers the copper.
- 3. Allow the reaction in the test tube to proceed for several minutes. Observe any changes.

Reaction B

- 4. Obtain a small chunk of zinc (Zn) metal. Place it into your test tube.
- 5. Carefully pour (or drip) hydrochloric acid (HCI) into the test tube until it covers the zinc.
- 6. Allow the reaction in the test tube to proceed for several minutes. Observe any changes.

Reaction C

- Add approximately 10 drops of lead nitrate (PbNO₃) into a small test tube.
- 8. Add 2 drops of potassium iodide (KI) to the test tube. Observe any changes.

Reaction D

- 9. Obtain a few chunks of copper (II) sulfate. Place them into your test tube.
- Light your Bunsen burner. Using a test tube holder, heat the test tube of copper (II) sulfate in the Bunsen burner flame for roughly 2 minutes. Observe any changes.

Reaction E

 Obtain a strip of magnesium. Carefully hold the magnesium in the Bunsen burner flame using tongs. When the reaction starts, do not look directly at the magnesium. Observe any changes.

Experiment 12 Types Of Chemical Reactions Lab

James McKeen Cattell, Jaques Cattell

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E-chemistry Iii Tm (science and Technology)' 2003 Ed., It's A Chemical World Sophia Vash, Frank Gasparro, Daniel Zibello, 2025-07-22 Chemistry is all around us it s a chemical world With this lab manual students and teachers alike will be prepared to conquer the wonders difficulties and fascinating fields of chemistry Featuring 17 carefully crafted chemistry laboratory experiments 14 lab demonstrations and five special experiments this book will provide you with a well rounded chemistry experience delving into multiple reaction types gas properties titrations calorimetry electrochemistry and more Finally it should be noted that these lab exercises have been well tested over the years These exercises can be used by students in either a regular course or an honors course **Experimental Organic Chemistry** Daniel R. Palleros, 2000-02-04 This cutting edge lab manual takes a multiscale approach presenting both micro semi micro and macroscale techniques The manual is easy to navigate with all relevant techniques found as they are needed Cutting edge subjects such as HPLC bioorganic chemistry multistep synthesis and more are presented in a clear and engaging **Nuclear Science Abstracts** ,1958 **Chemistry in the Laboratory** Jo A. Beran, 1996 This updated and revised fashion chemistry manual provides a well rounded understanding of concepts in the general chemistry laboratory Utilising visual aids experiments and equipment are explained and results and their pertinence discussed

Laboratory Manual to Accompany Introductory Chemistry Uno Kask, Otto W. Nitz, Martin G. Ondrus, Tammy J. Melton, 1993-07 This fifth edition of this laboratory manual emphasizes safety in the lab and discusses equipment requirements in the apparatus section at the beginning of each experiment It also features a revised art programme and explains the rational for each experiment

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