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3rd Edition

Chemical Bonds

2-1 If the isotope ^{32}S has 16 protons and 16 neutrons, how many protons and how many neutrons will the isotope ^{33}S have?

- 2-2** A. If 0.5 mole of glucose weighs 90 g, what is the molecular weight of glucose?
B. What is the concentration, in grams per liter (g/l), of a 0.25 M solution of glucose?
C. How many molecules are there in 1 mole of glucose?

2-3 Which of the following elements is LEAST abundant in living organisms?

- (a) Sulfur
(b) Carbon
(c) Oxygen
(d) Nitrogen
(e) Hydrogen

2-4 Your friend learns about Avogadro's number and thinks it is so huge that there may not even be a mole of living cells on Earth. You have recently heard that there are about 50 trillion (50×10^{12}) human cells in each adult human body, so you bet your friend \$5 that there is more than a mole of cells on Earth. Once you learn that each human contains more bacterial cells (in the digestive system) than human cells, you are sure that you have won the bet. The human population is now more than 6 billion (6×10^6). What calculation can you show your friend to convince him you are right?

2-5 Atoms form covalent bonds with each other by

- (a) sharing protons.
(b) sharing electrons.
(c) transferring electrons from one atom to the other.
(d) sharing neutrons.
(e) attraction of positive and negative charges.

2-6 A carbon atom contains six protons and six neutrons.

- A. What are its atomic number and atomic weight?
B. How many electrons does it have?
C. What is its valence? How does this affect carbon's chemical behavior?
D. Carbon with an atomic weight of 14 is radioactive. How does it differ in structure from

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**Bruce Alberts,Dennis Bray,Karen
Hopkin,Alexander D. Johnson,Julian
Lewis,Martin Raff,Keith Roberts,Peter
Walter**

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