

Answers to exam-style questions

Chapter 1

- 1 Solid [1] because both the melting point and boiling point are above 50°C. [1]
- 2 a Solids have particles that are very close together, are ordered and that vibrate about a fixed position. Answer: C [1]
- b Liquids have particles that are fairly close together, are irregularly arranged and that move slowly. Answer: B [1]
- c Gases have particles that are very far apart, are arranged randomly and that move at high speeds. Answer: D [1]
- d Particles that are very far apart must be in a gas. Gases have particles arranged randomly and moving at high speeds, which means they cannot be ordered or vibrating about a fixed position. Answer: A [1]
- 3 a Condensation [1]
- b Diffusion [1]
- c Freezing [1]
- d Evaporation [1]

- 4 a C [1]
- b E [1]
- c Separation: the particles get closer. [1]
Arrangement: particles are arranged in fixed positions. [1]
Motion: particles vibrate but do not move from place to place. [1]
Forces of attraction: particles are more strongly attracted. [1]

Chapter 2

- 1 a Iron(II) oxide has a formula: compound. [1]
- b Hematite is an ore: mixture. [1]
- c Iron appears in the Periodic Table: element. [1]
- d Stainless steel is an alloy: mixture. [1]
- e Air is a mixture of gases. [1]
- f Oxygen appears in the Periodic Table: element. [1]
- g Natural gas is mainly methane but also contains other gases: mixture. [1]
- h Methane has a formula: compound. [1]
- 2 a H = 2, S = 1, O = 4 [1]
- b C = 2, O = 1, H = 6 (5 from H₂ and 1 from OH) [1]
- c 5H₂O = 10 × H and 5 × O [1]
- Cu = 1, S = 1, H = 10, O = 5 + 4 = 9 [1]
- d Multiplying the brackets gives MgH₂O₄.
Mg = 1, N = 2, O = 6 [1]

- 3 a $\text{CaC}_2(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow \text{Ca}(\text{OH})_2(\text{aq}) + \text{C}_2\text{H}_2(\text{g})$ [1]
- b $2\text{KOH}(\text{aq}) + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{K}_2\text{SO}_4(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$ [1]
- c $\text{TiCl}_4(\text{l}) + 4\text{Na}(\text{s}) \rightarrow \text{Ti}(\text{s}) + 4\text{NaCl}(\text{s})$ [1]
- d $4\text{KO}_2(\text{s}) + 2\text{CO}_2(\text{g}) \rightarrow 2\text{K}_2\text{CO}_3(\text{g}) + 3\text{O}_2(\text{g})$ [1]
- e $2\text{Al}(\text{s}) + 6\text{HCl}(\text{aq}) \rightarrow 2\text{AlCl}_3 + 3\text{H}_2(\text{g})$ [1]
- 4 a Atomic number = number of protons = 20 [1]
- b Nucleon number = number of protons + number of neutrons = 20 + 21 = 41 [1]
- c 29 [1]
- d Number of neutrons = nucleon number – atomic number = 63 – 29 = 34 [1]
- e From the Periodic Table, element with atomic number 29 = copper [1]
- f 29 [1]
- g 29 + 34 = 63 [1]
- h From the Periodic Table, zinc has an atomic number of 30. [1]
- i 30 [1]
- j 30 + 35 = 65 [1]

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- Nucleus containing 3 protons [1] and 4 neutrons [1]
Three electrons outside the nucleus arranged 2,1 [1]
- 6 a Isotopes are atoms of the same element with the same number of protons but different numbers of neutrons. [1]

- b Consider 100 boron atoms:
20 boron atoms have mass of 10 relative mass units each
80 boron atoms have mass of 11 relative mass units each [1]
Total mass = (20 × 10) + (80 × 11) = 1080 [1]
Average mass = 1080 ÷ 100 = 10.80 [1]
Relative atomic mass (A_r) of boron = 10.80 [1]

- 7 Consider 100 copper atoms:
x = number of copper atoms with mass number 63
So (100 – x) = number of copper atoms with mass number 65 [1]
Total mass of 100 copper atoms = (63x) + 65(100 – x)
= 63x + 6500 – 65x [1]
= 6500 – 2x [1]

Foundation Chemistry Exam Style Questions With Answers

Carol Tear



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