

- ① For the reaction  $A+B \rightleftharpoons C+D$ ,  $\Delta H = +20 \text{ kJ/mol}$ . The activation energy of the forward reaction is  $85 \text{ kJ/mol}$ . The activation energy of the reverse reaction is.

①  $105 \text{ kJ/mol}$  ~~②  $65 \text{ kJ/mol}$~~  ③  $\frac{105 \text{ kJ/mol}}{2}$  ④  $\frac{65}{2} \text{ kJ/mol}$ .

Sol:- Activation energy of ~~Back~~ reverse reaction =  
Activation energy of forward reaction -  $\Delta H$ .  
 $\Rightarrow 85 - 20 = 65 \text{ kJ/mol}$ .

$$\therefore \Delta H = E_f - E_b$$

- ② The rate of a reaction is doubled of every  $10^\circ$  rise in temperature. The increase in the reaction rate as a result of temperature rise from  $10^\circ$  to  $100^\circ$  is.

① 112 ~~② 512~~ ③ 400 ④ 614.

Sol:- Rate of reaction  $= 2^n = 512$ .

|     |               |     |               |     |               |    |               |              |
|-----|---------------|-----|---------------|-----|---------------|----|---------------|--------------|
| 10  | $\rightarrow$ | 20  | $\rightarrow$ | 30  | $\rightarrow$ | 40 | $\rightarrow$ | 50           |
|     |               |     |               |     |               |    |               | $\downarrow$ |
| 100 | $\leftarrow$  | 90  | $\leftarrow$  | 80  | $\leftarrow$  | 70 | $\leftarrow$  | 60           |
| 512 |               | 256 |               | 128 |               | 64 |               | 32           |

- ③ Two reactions have rate constants  $K_1$  &  $K_2$  which are related as  $K_1 = 2K_2$ . If  $E_1$  &  $E_2$  are the activation energies of the first and the second reactions, the relation between them is.

①  $E_1 = 2E_2$  ②  $E_1 > E_2$  ~~③  $E_1 < E_2$~~  ④  $E_1 = \frac{1}{2} E_2$ .

Sol:- Fast reactions having low activation energy,  $E_1 < E_2$

- ④ In the reaction  $A+B \rightarrow \text{products}$ , the rate of reaction is doubled on doubling the concentration of A and is tripled when concentration of B increased by 9 times. The rate equation is.

# General Physics And Chemistry Objectives Questions And Answers

**Léo Pomar**



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