

# Chemistry Study Guide

### Periodic Table of the Elements

<div><div>H</div><div>1</div><div>1.008</div></div>																	<div><div>He</div><div>2</div><div>4.003</div></div>																						
<div><div>Li</div><div>3</div><div>6.941</div></div>	<div><div>Be</div><div>4</div><div>9.012</div></div>																	<div><div>B</div><div>5</div><div>10.811</div></div>	<div><div>C</div><div>6</div><div>12.011</div></div>	<div><div>N</div><div>7</div><div>14.007</div></div>	<div><div>O</div><div>8</div><div>15.999</div></div>	<div><div>F</div><div>9</div><div>18.998</div></div>	<div><div>Ne</div><div>10</div><div>20.180</div></div>																
<div><div>Na</div><div>11</div><div>22.990</div></div>	<div><div>Mg</div><div>12</div><div>24.305</div></div>																	<div><div>Al</div><div>13</div><div>26.982</div></div>	<div><div>Si</div><div>14</div><div>28.086</div></div>	<div><div>P</div><div>15</div><div>30.974</div></div>	<div><div>S</div><div>16</div><div>32.065</div></div>	<div><div>Cl</div><div>17</div><div>35.453</div></div>	<div><div>Ar</div><div>18</div><div>39.948</div></div>																
<div><div>K</div><div>19</div><div>39.098</div></div>	<div><div>Ca</div><div>20</div><div>40.078</div></div>	<div><div>Sc</div><div>21</div><div>44.956</div></div>	<div><div>Ti</div><div>22</div><div>47.883</div></div>	<div><div>V</div><div>23</div><div>50.942</div></div>	<div><div>Cr</div><div>24</div><div>51.996</div></div>	<div><div>Mn</div><div>25</div><div>54.938</div></div>	<div><div>Fe</div><div>26</div><div>55.845</div></div>	<div><div>Co</div><div>27</div><div>58.933</div></div>	<div><div>Ni</div><div>28</div><div>58.693</div></div>	<div><div>Cu</div><div>29</div><div>63.546</div></div>	<div><div>Zn</div><div>30</div><div>65.38</div></div>	<div><div>Ga</div><div>31</div><div>69.723</div></div>	<div><div>Ge</div><div>32</div><div>72.630</div></div>	<div><div>As</div><div>33</div><div>74.922</div></div>	<div><div>Se</div><div>34</div><div>78.96</div></div>	<div><div>Br</div><div>35</div><div>79.904</div></div>	<div><div>Kr</div><div>36</div><div>83.80</div></div>																						
<div><div>Rb</div><div>37</div><div>85.468</div></div>	<div><div>Sr</div><div>38</div><div>87.62</div></div>	<div><div>Y</div><div>39</div><div>88.906</div></div>	<div><div>Zr</div><div>40</div><div>91.224</div></div>	<div><div>Nb</div><div>41</div><div>92.906</div></div>	<div><div>Mo</div><div>42</div><div>95.94</div></div>	<div><div>Tc</div><div>43</div><div>98.906</div></div>	<div><div>Ru</div><div>44</div><div>101.07</div></div>	<div><div>Rh</div><div>45</div><div>102.91</div></div>	<div><div>Pd</div><div>46</div><div>106.42</div></div>	<div><div>Ag</div><div>47</div><div>107.87</div></div>	<div><div>Cd</div><div>48</div><div>112.41</div></div>	<div><div>In</div><div>49</div><div>114.82</div></div>	<div><div>Sn</div><div>50</div><div>118.71</div></div>	<div><div>Sb</div><div>51</div><div>121.76</div></div>	<div><div>Te</div><div>52</div><div>127.6</div></div>	<div><div>I</div><div>53</div><div>126.91</div></div>	<div><div>Xe</div><div>54</div><div>131.29</div></div>																						
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<div><div>Fr</div><div>87</div><div>223</div></div>	<div><div>Ra</div><div>88</div><div>226</div></div>																	<div><div>Uut</div><div>113</div><div>288</div></div>	<div><div>Fl</div><div>114</div><div>289</div></div>	<div><div>Uup</div><div>115</div><div>288</div></div>	<div><div>Lv</div><div>116</div><div>293</div></div>	<div><div>Uus</div><div>117</div><div>294</div></div>	<div><div>Uuo</div><div>118</div><div>294</div></div>																
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## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Genral Chemistry Study Guide

**Raymond E. Davis**



## General Chemistry Study Guide:

**Study Guide for Ebbing/Gammons General Chemistry** Darrell Ebbing, Steven D. Gammon, 2016-02-05 The perfect way to prepare for exams build problem solving skills and get the grade you want This guide reinforces your understanding of major concepts learning objectives and key terms presented in your text and further develops your problem solving skills Each chapter features both a diagnostic pre test and post test additional practice problems and their worked out solutions and cumulative unit exams

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