

Reaction	Equilibrium Constant	Reaction	Equilibrium Constant
<p>Reaction 1: $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$</p> <p>Equilibrium Constant: K_1</p> 		<p>Reaction 2: $2\text{NH}_3 \rightleftharpoons \text{N}_2 + 3\text{H}_2$</p> <p>Equilibrium Constant: K_2</p> 	
<p>Reaction 3: $\text{N}_2 + \text{O}_2 \rightleftharpoons 2\text{NO}$</p> <p>Equilibrium Constant: K_3</p> 		<p>Reaction 4: $2\text{NO} \rightleftharpoons \text{N}_2 + \text{O}_2$</p> <p>Equilibrium Constant: K_4</p> 	
<p>Reaction 5: $\text{N}_2 + 2\text{O}_2 \rightleftharpoons 2\text{NO}_2$</p> <p>Equilibrium Constant: K_5</p> 		<p>Reaction 6: $2\text{NO}_2 \rightleftharpoons \text{N}_2 + 2\text{O}_2$</p> <p>Equilibrium Constant: K_6</p> 	
<p>Reaction 7: $\text{N}_2 + 2\text{O}_2 \rightleftharpoons 2\text{NO}_2$</p> <p>Equilibrium Constant: K_7</p> 		<p>Reaction 8: $2\text{NO}_2 \rightleftharpoons \text{N}_2 + 2\text{O}_2$</p> <p>Equilibrium Constant: K_8</p> 	
<p>Reaction 9: $\text{N}_2 + 2\text{O}_2 \rightleftharpoons 2\text{NO}_2$</p> <p>Equilibrium Constant: K_9</p> 		<p>Reaction 10: $2\text{NO}_2 \rightleftharpoons \text{N}_2 + 2\text{O}_2$</p> <p>Equilibrium Constant: K_{10}</p> 	

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