EXPERIMENT 19

An Analysis of Commercial Antacids







OBJECTIVES:

Determine the per gram neutralizing ability of a commercial antacid Determine the per gram neutralizing ability of sodium bicarbonate Compare the effectiveness of several antacids.

INTRODUCTION:

"How do you spell relief?"

The parietal cells in the stomach secrete sufficient hydrochloric acid to normally keep the acidity at about 0.10 M. The flow of HCl increases when food enters the stomach and so acidity increases when you eat or drink. If overstimulation of gastric juices builds the concentration up too much, heartburn or indigestion often results. Sodium bicarbonate (NaHCO₃ in baking soda) has historically been used to try to relieve this condition. Today we have a number of commercial products which purport to neutralize (or buffer) excess stomach acid. Television commercials claiming greater effectiveness in neutralizing acid are familiar. Rolaids claims to consume 47 times its weight in excess stomach acid. Tums claims to have one third more neutralizing power than other antacids. In this lab, we will compare acid-neutralizing capabilities for ourselves.

Several types of commercial antacids will be tested and the results will be shared in class. Keep in mind that tablets may also contain flavors, sweeteners, binders, fillers, antifoam agents, pain relievers, etc. In this experiment, the tablets will be analyzed only for their ability to neutralize acids. The active ingredient in antacids varies with the brand. The table below lists active ingredients in several brands.

COMMERCIAL ANTACIDS' MAJOR ACTIVE INGREDIENTS

Brand	Active Agent
Milk of Magnesia	Mg(OH)-
Rolaids	NaAl(OH),CO1
Turns	CaCO ₃
Alka-Seltzer II	NaHCO ₁ and KHCO ₂
Maslox	Mg(OH), and Al(OH),
Gaviscon	Al(OH)

The H (actually H₁O) from acids are neutralized by these agents as illustrated below:

insoluble carbonates:
$$CaCO_3(s) + 2 H^*(aq) \longrightarrow Ca^{2*}(aq) + H_2O(l) + CO_2(g)$$

Effective Commercial Antacids Experiment

Phyllis J. Perry

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