

Acid Base Strength Table

TABLE 8.2 Some Acids and Their Conjugate Bases, in Decreasing Order of Acid Strength

Acid		Conjugate Base		
 Strong Acids	HI	Hydroiodic acid	I^-	Iodide ion
	H_2SO_4	Sulfuric acid	HSO_4^-	Hydrogen sulfate ion
	HCl	Hydrochloric acid	Cl^-	Chloride ion
	HNO_3	Nitric acid	NO_3^-	Nitrate ion
	H_3O^+	Hydronium ion	H_2O	Water
	HSO_4^-	Hydrogen sulfate ion	SO_4^{2-}	Sulfate ion
	H_3PO_4	Phosphoric acid	$H_2PO_4^-$	Dihydrogen phosphate ion
	CH_3COOH	Acetic acid	CH_3COO^-	Acetate ion
	H_2CO_3	Carbonic acid	HCO_3^-	Bicarbonate ion
	H_2S	Hydrogen sulfide	HS^-	Hydrogen sulfide ion
	$(H_2PO_4)^-$	Dihydrogen phosphate ion	HPO_4^{2-}	Hydrogen phosphate ion
	NH_4^+	Ammonium ion	NH_3	Ammonia
	C_6H_5OH	Phenol	$C_6H_5O^-$	Phenoxide ion
	HCO_3^-	Bicarbonate ion	CO_3^{2-}	Carbonate ion
	HPO_4^{2-}	Hydrogen phosphate ion	PO_4^{3-}	Phosphate ion
	H_2O	Water	OH^-	Hydroxide ion
	C_2H_5OH	Ethanol	$C_2H_5O^-$	Ethoxide ion
	NH_3	Ammonia	NH_2^-	Amide ion
	 Weak Bases			

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Acids and Bases: Their Quantitative Behaviour Ronald Percy Bell,1969 Acid-base Equilibria Edward Jasper King,1965 Band 4 **Acids and Bases** Brian G. Cox,2013-01-31 Acids and bases are ubiquitous in chemistry Our understanding of them however is

dominated by their behaviour in water Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of substrates This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents The behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented Fundamental background material is provided in the initial chapters quantitative aspects of acid base equilibria including definitions and relationships between solution pH and species distribution the influence of molecular structure on acid strengths and acidity in aqueous solution Solvent properties are reviewed along with the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial Experimental methods for determining dissociation constants are described in detail In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described *Solid Acids and Bases* Kozo Tanabe,2012-12-02 Solid Acids and Bases Their Catalytic Properties reviews developments in the studies of acidic and basic properties of solids including the efficacy and special characteristics of solid acid and base catalysts This book discusses the determination of basic and acidic properties on solid surfaces and relationship between acid strength and acid amount The structure and acid base properties of mixed metal oxides and correlation between acid base properties and catalytic activity and selectivity are also deliberated This publication is useful to professional chemists and graduate students in the fields of

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